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a) Limewater is the common name for a diluted solution of calcium hydroxide
c) Range 1-3 d) Strong Base is NaOH and Weak Base is NH_4OH e) A hydrated salt is copper sulphate crystals ($\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$). f) A universal indicator is a mixed indicator of organic chemicals which not only shows whether the given solution is acidic or basic, but also shows the approximate pH values by giving a wide ...

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in the blanks (a) The most commonly used natural indicator is _____ and is extracted from _____ ...

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1. acids 2. citric acid 3. Base 4. calcium sulphate hemihydrate, $\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$ 5. yellow colour, pink colour 6. Hydrochloric Acid (HCl) 7. Acetic Acid (CH_3COOH) 8. Sodium Hydroxide (NaOH) 9. Ammonium Hydroxide (NH_4OH) 10. sodium carbonate, $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$, permanent hardness 11. PH 12. Acid 13. Base 14. Neutral 15. decreases 16 ...

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Acid + Base \rightarrow Salt + Water Since, the reaction between acid and base both neutralize each other, hence, it is also known as Neutralization Reaction. Examples: Sodium chloride and water are formed when hydrochloric acid reacts with sodium hydroxide (a strong base).

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Neutralization is the reaction between acid and base to form salt and water only. From ionic equation, we know that the reaction only involves H^+ ions from acids with OH^- ions from alkali to form water. E.g. $NaOH + H_2SO_4$ forms $Na_2SO_4 + H_2O$

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10. Explain in terms of the partial charge on hydrogen why $NaOH$ is a base, $HClO$ is a weak acid and $HClO_4$ is a strong acid. 11. Why is HCl a strong acid and $HClO$ a weak acid? 12. Why are HCl and $HClO_4$ both strong acids? 13. For each of the reactions below, classify the reactants as an acid or a base and the products as the conjugate acid or ...

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