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Compound 1: 21.8 g C and 58.2 g O
(80.0 - 21.8 = mass O) Compound 2:
34.3 g C and 45.7 g O (80.0 - 34.3 =
mass O) The mass of carbon that
combines with 1.0 g of oxygen is:
Compound 1: Compound 2: 21.8 g C
58.2 g O 34.3 g C 45.7 g O = 0.375 g C/g
O = 0.751 g C/g O The ratio of the
masses of carbon that combine with 1 g
of oxygen is 0.751 = 0 ...

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8.71 Bonds broken: 1 C-N (305 kJ/mol)
Bonds formed: 1 C-C (347 kJ/mol) $\Delta H =$
 $\sum D_{\text{broken}} - \sum D_{\text{formed}}$, $\Delta H = 305 - 347 =$
 -42 kJ Note: sometimes some of the
bonds remain the same between
reactants and products. To save time,
only break and form bonds that are
involved in the reaction. 8.73 Bonds
broken: Bonds formed:

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